

**VITEEE – 2023**  
**CHEMISTRY**

**1. Physical Chemistry**

*Atomic Structure* - Bohr's atomic model-Sommerfeld's extension of atomic structure; Electronic configuration and Quantum numbers; Shapes of s, p, d, f orbitals - Pauli's exclusion principle - Hund's Rule of maximum multiplicity- Aufbau principle. Emission and absorption spectra, line and band spectra; Hydrogen spectrum – Lyman, Balmer, Paschen, Brackett and Pfund series; de Broglie's theory; Heisenberg's uncertainty principle – wave nature of electron – Schrodinger wave equation (No derivation). Eigen values and eigen functions. Chemical bonding and hybridization of atomic orbitals involving s, p and d orbitals.

*Thermodynamics, Chemical Equilibrium and Chemical Kinetics* - I and II Laws of thermodynamics – spontaneous and non-spontaneous processes, entropy, Gibb's free energy – Standard Gibbs free energy change ( $\Delta G^0$ ) and chemical equilibrium – significance of entropy. Rate of a chemical reaction, factors affecting rates of reaction: concentration, temperature, pressure and catalyst; Law of mass action – Le Chatelier's principle, applications of chemical equilibrium. Rate expression, order, and molecularity of reactions, zero order, first order and pseudo first order reaction – half-life period. Determination of rate constant and order of reaction. Temperature dependence of rate constant – Arrhenius equation, activation energy and its calculation; elementary concept of collision theory of bimolecular gaseous reactions.

*Solutions* - Colligative properties of dilute solutions; Different methods for expressing the concentration of solution - molality, molarity, mole fraction, percentage, the vapour pressure of solutions and Raoult's Law - Ideal and non-ideal solutions, vapour pressure - composition, plots for ideal and non-ideal solutions

**2. Inorganic and Material Chemistry**

*The s-block elements* – properties and chemical reactivity of alkali and alkaline earth metals

*The p-block elements* – Phosphorous compounds:  $\text{PCl}_3$ ,  $\text{PCl}_5$  – Oxides, Hydrogen halides, Inter-halogen compounds and Xenon fluoride compounds

*General characteristics of d – block elements* – Electronic Configuration – Oxidation states of first row transition elements and their colours. Occurrence and principles of extraction: Copper, Silver, Gold and Zinc. Preparation and properties of  $\text{CuSO}_4$ ,  $\text{AgNO}_3$  and  $\text{K}_2\text{Cr}_2\text{O}_7$ .

*Lanthanides* – Introduction, electronic configuration, general characteristics, oxidation state – lanthanide contraction, uses, brief comparison of Lanthanides and Actinides

*Introduction to coordination chemistry* - IUPAC nomenclature of mononuclear coordination compounds; Isomerism, Geometrical isomerism in 4-coordinate, 6-coordinate complexes. Theories on coordination compounds – Werner's theory (brief), Valence Bond theory. Uses of coordination compounds. Bioinorganic compounds (Haemoglobin and chlorophyll).

*Solid-State Chemistry* - Lattice – unit cell, systems, types of crystals, packing in solids; Ionic crystals – Imperfections in solids – point defects, X-Ray diffraction – Electrical Property, Amorphous solids (elementary ideas only)

*Surface Chemistry* - Adsorption- physisorption and chemisorption; Catalysis – homogeneous and heterogeneous catalysis

### 3. Analytical Chemistry

*Electrochemistry* - Redox reactions; Theory of electrical conductance; metallic and electrolytic conductance. Faraday's laws – theory of strong electrolytes – Specific resistance, specific conductance, equivalent and molar conductance – Variation of conductance with dilution – Kohlrausch's Law – Ionic product of water, pH, and pH– buffer solutions – use of pH values. Cells – Electrodes and electrode potentials – construction of cell, EMF values and standard electrode potentials, Nernst equation and its application to chemical cells. Relation between Gibbs energy change and EMF of a cell, dry cell, electrolytic cells and Galvanic cells; lead accumulator; Fuel cells, Corrosion, and its prevention.

*Environmental Chemistry* - Environmental pollution - Atmospheric, water and soil.

### 4. Basic Principles of Organic Chemistry

*Carbon* – tetravalency, hybridization; Classification of organic compounds – functional groups; Homologous series; Nomenclature (IUPAC); Homolytic and heterolytic bond cleavage; carbocations, carbanions and free radicals; electrophiles and nucleophiles; Inductive effect, electromeric effect, resonance and hyperconjugation.

*Common organic reactions* - Substitution, addition, elimination and rearrangement

*Isomerism in Organic Compounds*: Definition, Classification – structural isomerism, stereo isomerism – geometrical and optical isomerism. Optical activity - chirality – compounds containing chiral centres – R, S notation, D, L notation.

*Detection of the functional groups in organic compounds*: Hydroxyl (alcoholic and phenolic), carbonyl (aldehyde and ketones) carboxyl and amino groups.

### 5. Properties and Chemistry of Functionalized Organic Compounds

*Alcohols and Ethers* - Nomenclature of alcohols – Classification of alcohols - distinction between 1°, 2° and 3° alcohols – General methods of preparation of primary alcohols, properties. Methods of preparation of dihydric alcohols: Glycol – Properties – Uses. Methods of preparation of trihydric alcohols - Properties – Uses. Aromatic alcohols – preparation and properties of phenols and benzyl alcohol; Ethers – Nomenclature of ethers – general methods of preparation of aliphatic ethers - Properties – Uses. Aromatic ethers – Preparation of Anisole – Uses

*Carbonyl Compounds* - Nomenclature of carbonyl compounds – Comparison of aldehydes and ketones. General methods of preparation of aldehydes – Properties – Uses. Aromatic aldehydes – Preparation of benzaldehyde – Properties and Uses. Ketones – general methods of preparation of aliphatic ketones (acetone) – Properties – Uses. Aromatic ketones – preparation of acetophenone – Properties – Uses, preparation of benzophenone – Properties. Name reactions; Clemmensen reduction, Wolff – Kishner reduction, Cannizzaro reaction, Claisen Schmidt reaction, Benzoin Condensation, Aldol Condensation. Preparation and applications of Grignard reagents.

*Carboxylic Acids and their derivatives* - Nomenclature – Preparation of aliphatic monocarboxylic acids – formic acid – Properties – Uses. Monohydroxy mono carboxylic acids; Lactic acid – Synthesis of lactic acid. Aliphatic dicarboxylic acids; Preparation of oxalic and succinic acids. Aromatic acids: Benzoic and Salicylic acids – Properties – Uses. Derivatives of carboxylic acids; acetyl chloride ( $\text{CH}_3\text{COCl}$ ) – Preparation – Properties – Uses. Preparation of acetamide, Properties – acetic anhydride – Preparation, Properties. Preparation of esters – methyl acetate – Properties

## 6. Organic Nitrogen Compounds

*Organic Nitrogen Compounds* - Aliphatic nitro compounds – Preparation of aliphatic nitroalkanes – Properties – Uses. Aromatic nitro compounds – Preparation – Properties – Uses. Distinction between aliphatic and aromatic nitro compounds. Amines; aliphatic amines – General methods of preparation – Properties – Distinction between 1°, 2° and 3° amines. Aromatic amines – Synthesis of benzylamine – Properties, Aniline – Preparation – Properties – Uses. Differences between aliphatic and aromatic amines. Aliphatic nitriles – Preparation – properties – Uses. Diazonium salts – Preparation of benzene diazonium chloride – Properties.

## 7. Biomolecules and Polymers

*Carbohydrates* – Distinction between sugars and non-sugars, structural formulae of glucose, fructose, and sucrose, with their linkages, invert sugar – definition, examples of oligo and polysaccharides

*Amino acids and Proteins* – Classification of amino acids with examples, Peptides - properties of peptide bond; Proteins - primary, secondary, tertiary and quaternary structure (qualitative idea only), denaturation of proteins, enzymes

*Lipids* - Definition, classification with examples, difference between fats, oils, and waxes.

*Nucleic acids* – Chemical constitution of DNA and RNA

*Polymers* - Classification – Natural and synthetic, methods of polymerization (addition and condensation), copolymerization. Some important polymers: natural and synthetic like polythene, nylon, polyesters, Bakelite, rubber. Biodegradable and non-biodegradable polymers.

